

عنوان درس

فيزيک حالت جامد ۲ Solid-state physics 2







فيزيك حالت جامد چيست؟

- Solid-state physics is the study of rigid <u>matter</u>, or <u>solids</u>, through methods such as <u>quantum mechanics</u>, <u>crystallography</u>, <u>electromagnetism</u>, and <u>metallurgy</u>.
- ✤ It is the largest branch of <u>condensed matter physics</u>.
- Solid-state physics studies how the large-scale properties of solid materials result from their <u>atomic</u>-scale properties.
- Thus, solid-state physics forms a theoretical basis of <u>materials</u> <u>science</u>.
- It also has direct applications, for example in the technology of <u>transistors</u> and <u>semiconductors</u>.

https://en.wikipedia.org/wiki/Solid-state\_physics

## حوزه بحث فيزيك حالت جامد

- Solid materials are formed from densely packed atoms, which interact intensely.
- These interactions produce the mechanical (e.g. <u>hardness</u> and <u>elasticity</u>), <u>thermal</u>, <u>electrical</u>, <u>magnetic</u> and <u>optical</u> properties of solids.
- Depending on the material involved and the conditions in which it was formed, the atoms may be arranged in a regular, geometric pattern (crystalline solids, which include metals and ordinary water ice) or irregularly (an amorphous solid such as common window glass).

حوزه بحث فيزيك حالت جامد

- The bulk of solid-state physics, as a general theory, is focused on <u>crystals</u>.
- Primarily, this is because the <u>periodicity</u> of <u>atoms</u> in a crystal its defining characteristic facilitates mathematical modeling.
- Likewise, crystalline materials often have <u>electrical</u>, <u>magnetic</u>, <u>optical</u>, or <u>mechanical</u> properties that can be exploited for <u>engineering</u> purposes.

حوزه بحث فيزيك حالت جامد

The forces between the atoms in a crystal can take a variety of forms.

- ✤ For example, in a crystal of <u>sodium chloride</u> (common salt), the crystal is made up of <u>ionic sodium</u> and <u>chlorine</u>, and held together with <u>ionic bonds</u>.
- ✤In others, the atoms share <u>electrons</u> and form <u>covalent bonds</u>. In metals, electrons are shared amongst the whole crystal in <u>metallic bonding</u>.
- ✤Finally, the noble gases do not undergo any of these types of bonding. In solid form, the noble gases are held together with <u>van der Waals forces</u> resulting from the polarisation of the electronic charge cloud on each atom.
- The differences between the types of solid result from the differences between their bonding.